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#### **Key indicators**

Single-crystal X-ray study T = 293 K Mean  $\sigma$ (C–C) = 0.004 Å R factor = 0.021 wR factor = 0.055 Data-to-parameter ratio = 12.5

For details of how these key indicators were automatically derived from the article, see http://journals.iucr.org/e.

# Poly[ $\mu_2$ -4,4'-bipyridine-di- $\mu_2$ -imidazolidocadmium(II)]

The title compound,  $[Cd(C_{10}H_8N_2)(C_3H_4N_2)_2]_n$ , is a new neutral centrosymmetric three-dimensional metal-organic framework structure. The metal atoms occupy inversion centres and the bipyridine ligands also possess  $C_i$  symmetry. The metal atoms are bridged by four imidazole ligands and two bipyridine units.

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## Comment

The construction of metal-organic frameworks (MOFs) is currently receiving considerable attention owing to their potential properties as functional solid materials, as well as their facinating framework structures (Evans & Lin, 2002; Lu et al., 2001; Li et al., 1999). New materials spur new technologies (Snurr et al., 2004), and these compounds can be potentially useful in gas storage, as molecular sieves, in sizeand shape-selective catalysis, in magnetism, in optoelectronic devices, and so on (Eddaoudi et al., 2002; Kitaura et al., 2003; Moulton & Zaworotko, 2001; Millange et al., 2002). These MOF structures can be rapidly, accurately and efficiently synthesized from relatively simple subunits, where the metal ions, multidentate organic ligands and coordinating bonds are the parameters for directing the self-assembly process. Our studies on the complexation of cadmium(II) salts with 4,4'bipyridine (bpy) and imidazole has led to the formation of the title compound, (I), which has proved to have a polymeric centrosymmetric three-dimensional structure.



A segment of the polymeric structure of compound (I) is illustrated in Fig. 1, and selected bond distances and angles are given in Table 1. The  $Cd^{II}$  atom lies on an inversion centre and the bipyridine ligand possesses  $C_i$  symmetry. The metal atoms

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Figure 1

A view of part of the polymeric structure of compound (I), showing the atom-labelling scheme and with displacement ellipsoids drawn at the 50% probability level. [Symmetry codes: (A) -x, -y, -z + 1; (B)  $x, -y + \frac{1}{2}, z - \frac{1}{2}$ ; (C)  $-x, y - \frac{1}{2}, -z + \frac{3}{2}$ .]

are bridged by four imidazole ligands and two bipyridine units. The Cd $-N_{imidazole}$  bond lengths vary from 2.324 (2) to 2.342 (2) Å and the Cd $-N_{bipyridine}$  distances are 2.513 (2) Å. The *cis*-N-Cd-N bond angles range from 87.96 (7) to 92.04 (7)° (Table 1). The N atoms coordinate to the Cd<sup>II</sup> centres in the monodentate mode (Fig. 1). On the basis of this connectivity, an interesting three-dimensional structure is built up, as shown in Fig. 2.

## Experimental

Compound (I) was obtained by a hydrothermal reaction. The starting materials,  $Cd(OAc)_2 \cdot 4H_2O$  (0.533 g, 2 mmol), imidazole (0.136 g, 2 mmol), 4,4'-bipyridine (0.1922 g, 1 mmol), and deionized water (15 ml) were mixed and the pH of the solution adjusted to 6.5 by the addition of sodium hydroxide (*ca* 0.08 g). The resulting suspension was stirred for 1 h, then sealed in a 40 ml stainless-steel bomb with a Teflon liner and heated at 458 K for 4 d. The autoclave was then cooled at 6 K h<sup>-1</sup> to 376 K. This temperature was maintained for 12 h, and then the system was cooled to room temperature at the same cooling rate. Colourless crystals of (I) were obtained and washed with water.

### Crystal data

3682 measured reflections

$[Cd(C_{10}H_8N_2)(C_3H_4N_2)_2]$	$D_x = 1.788 \text{ Mg m}^{-3}$
$M_r = 402.73$	Mo $K\alpha$ radiation
Monoclinic, $P2_1/c$	Cell parameters from 2578
a = 9.2193 (13)  Å	reflections
b = 7.9209 (12)  Å	$\theta = 3.3-26.4^{\circ}$
c = 10.2436 (15)  Å	$\mu = 1.47 \text{ mm}^{-1}$
$\beta = 90.034 \ (2)^{\circ}$	T = 293 (2) K
$V = 748.04 (19) \text{ Å}^3$	Block, colourless
Z = 2	$0.20$ $\times$ 0.18 $\times$ 0.10 mm
Data collection	
Bruker SMART 1000 CCD area-	1323 independent reflections
detector diffractometer	1150 reflections with $I > 2\sigma(I)$
$\varphi$ and $\omega$ scans	$R_{\rm int} = 0.025$
Absorption correction: multi-scan	$\theta_{\rm max} = 25.1^{\circ}$
(SADABS; Bruker, 2000)	$h = -10 \rightarrow 9$
$T_{\min} = 0.560, T_{\max} = 0.863$	$k = -7 \rightarrow 9$

 $l = -7 \rightarrow 12$ 





Refinement

 $\begin{array}{ll} \mbox{Refinement on } F^2 & w = 1/[\sigma^2(F_o^2) + (0.0205P)^2 \\ R[F^2 > 2\sigma(F^2)] = 0.021 & where \ P = (F_o^2 + 2F_c^2)/3 \\ wn (F^2) = 0.055 & where \ P = (F_o^2 + 2F_c^2)/3 \\ S = 1.14 & (\Delta/\sigma)_{max} = 0.003 \\ 1323 \ reflections & \Delta\rho_{max} = 0.26 \ e \ {\rm \AA}^{-3} \\ 106 \ parameters & \Delta\rho_{min} = -0.67 \ e \ {\rm \AA}^{-3} \\ \mbox{H-atom parameters constrained} \\ \end{array}$ 

Table 1 Selected geometric parameters (Å, °).

Cd1-N1 Cd1-N2 <sup>i</sup>	2.343 (2) 2.324 (2)	Cd1-N3	2.513 (2)
$N1^{ii} - Cd1 - N1$ $N2^{i} - Cd1 - N2^{iii}$ $N3^{ii} - Cd1 - N3$ $N2^{i} - Cd1 - N1^{ii}$ $N2^{iii} - Cd1 - N1^{ii}$	180 180 180 87.96 (7) 92.04 (7)	$N2^{i}-Cd1-N3^{ii}$ $N2^{iii}-Cd1-N3^{ii}$ $N1^{ii}-Cd1-N3^{ii}$ $N1-Cd1-N3^{ii}$	90.19 (7) 89.81 (7) 89.25 (7) 90.75 (7)

Symmetry codes: (i)  $x, -y + \frac{1}{2}, z - \frac{1}{2}$ ; (ii) -x, -y, -z + 1; (iii)  $-x, y - \frac{1}{2}, -z + \frac{3}{2}$ .

The H atoms were placed in calculated positions and constrained to ride on their parent atoms, with C-H distances of 0.93 Å and  $U_{iso}(H) = 1.2U_{eq}(C)$ .

Data collection: *SMART* (Bruker, 2000); cell refinement: *SAINT* (Bruker, 2000); data reduction: *SAINT*; program(s) used to solve structure: *SHELXTL* (Bruker, 2000); program(s) used to refine structure: *SHELXTL*; molecular graphics: *PLATON* (Spek, 2003); software used to prepare material for publication: *SHELXTL*.

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#### References

Eddaoudi, M., Kim, J., Rosi, N., Vodak, D., Wachter, J., O'Keeffe, M. & Yaghi, O. M. (2002). *Science*, **295**, 469–472.

- Evans, O. R. & Lin, W. (2002). Acc. Chem. Res. 35, 511-522.
- Bruker (2000). SMART (Version 5.6), SAINT (Version 6.1), SADABS (Version 2.01) and SHELXTL (Version 6.1). Bruker AXS Inc., Madison, Wisconsin, USA.
- Kitaura, R., Seki, K., Akiyama, G. & Kitagawa, S. (2003). Angew. Chem. Int. Ed. 42, 428–431.

- Li, H., Eddaoudi, M., O'Keeffe, M. & Yaghi, O. M. (1999). *Nature (London)*, **402**, 276–279.
- Lu, J., Mondal, A. & Zaworotko, M. J. (2001). Angew. Chem. Int. Ed. 113, 2171–2174.

Millange, F., Serre, C. & Ferey, G. (2002). *Chem. Commun.* pp. 822–823. Moulton, B. & Zaworotko, M. J. (2001). *Chem. Rev.* **101**, 1629–1658. Snurr, R. Q., Hupp, J. T. & Nguyen, S. T. (2004). *AIChE*, **50**, 1090–1095. Spek, A. L. (2003). *J. Appl. Cryst.* **36**, 7–13.